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|  | **Vocabulary Term** | **Definition** | **Illustration or Used in a Sentence** |
| 1 | Precipitate | Solid that occurs from the mixing of two water solutions of ionic compounds. |  |
| 2 | Exothermic reaction | A chemical reaction that RELEASES heat |  |
| 3 | Endothermic reaction | A chemical reaction that ABSORBS heat |  |
| 4 | Law of Conservation of Mass | A natural law describing the fact that matter is neither created nor destroyed. A chemical reaction will have the same mass in its reactants as it does in its products |  |
| 5 | Single replacement | A chemical reaction in which an uncombined element replaces another element in a compound.  Example A + BX 🡪 AX + B |  |
| 6 | Double replacement | A chemical reaction in which elements from two different compounds trade places.  Example AX + BY 🡪 AY + BX |  |
| 7 | Synthesis (Direct Combination) | A chemical reaction in which two or more elements or compounds combine and form one product.  Example A + B 🡪 AB |  |
| 8 | Decomposition | A chemical reaction in which a single complex compound is broken down into two or more products  Example AB 🡪 A + B |  |
| 9 | Combustion (burning) | A chemical reaction in which heat and light are given off when oxygen combines with another compound (the fuel) to produce water (H2O) and carbon dioxide (CO2).  Example **C10H8 + 12 O2 ---> 10 CO2 + 4 H2O** |  |

**Vocabulary – Chemical Reactions #2**