Name

Date

Period

Review of Mixtures

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| **Station 1- Aluminum Staples**  Formula- Al   1. Is this station a mixture or a pure substance? 2. Is this an example of an element or a compound? 3. What state of matter is Station 1? | **Station 2- Empty Soda Can**  Formula- Al + O2 + Mn   1. Is the soda can a mixture or pure substance? 2. How do you know? 3. What state of matter is the can at Station 2? |
| **Station 3- Snack Mix**   1. Is this station a mixture or a pure substance? 2. Which type of mixture is the snack mix? 3. If the snack mix were to be broken in tiny pieces would that be a physical or chemical change? | **Station 4- Rotting Fruit**   1. Is this a physical or chemical change? 2. How do you know? 3. What state of matter is Station 4? |
| **Station 5- Rusty Metal**  Formula- 4Fe + 3O2 + 6H2O -> 4Fe(OH)3   1. Is this station a mixture or a pure substance? 2. How do you know? 3. When the metal rusted was that a physical or   chemical change?   1. What state of matter is Station 5? | **Station 6- Salt Water**  Formula- NaCl + H2O   1. Is the salt a mixture or pure substance? 2. Is the salt water a heterogeneous or homogeneous mixture? 3. What state of matter is Station 6? 4. How can we get the salt back? |

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| **Station 7- One round ball (use your imagination)**   1. Is this an example of an atom or molecule? 2. How do you know? 3. What state of matter is the ball at Station 7? | **Station 8- Dried Pasta and Beans**   1. Is the pasta and the beans an example of a mixture or pure substance? 2. Is the pasta and bean mixture a heterogeneous or homogeneous mixture? 3. What state of matter is Station 8? |
| **Station 9- Sand**  Formula- **SiO2 + CaSiO4 + Al2O3**   1. Is the sand an example of a mixture or pure substance? 2. Is the sand a heterogeneous or homogeneous mixture? 3. What state of matter is Station 9? | **Station 10- Sugar Water**  Formula- H2O + C12H22O11   1. Is the sugar water an example of a mixture or pure substance? 2. Is the sugar water a heterogeneous or homogeneous mixture? 3. What state of matter is Station 10? 4. How can I turn water into a gas? |
| ***Extra Credit***  **Station 11- Melted Crayons**   1. Are these melted crayons a mixture or pure substance? 2. Is this an example of a chemical or physical change? 3. What could I do with the crayons to make a chemical change occur? | ***Extra Credit***  **Station 12- Air in a Balloon**  Formula- 02   1. Is the air inside of balloon a mixture or a pure substance? 2. Is the air inside an element or a molecule? 3. What state of matter is the air? |

13) Look at the above Stations with “formulas”, using your textbook’s periodic table identify what each chemical abbreviation stands for. (Example- Mg= Magnesium) \*\*Each time you see a capital letter it represents a new element.