Name

 Date

 Period

 Review of Mixtures

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| **Station 1- Aluminum Staples** Formula- Al1. Is this station a mixture or a pure substance?
2. Is this an example of an element or a compound?
3. What state of matter is Station 1?
 | **Station 2- Empty Soda Can** Formula- Al + O2 + Mn1. Is the soda can a mixture or pure substance?
2. How do you know?
3. What state of matter is the can at Station 2?
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| **Station 3- Snack Mix**1. Is this station a mixture or a pure substance?
2. Which type of mixture is the snack mix?
3. If the snack mix were to be broken in tiny pieces would that be a physical or chemical change?
 | **Station 4- Rotting Fruit**1. Is this a physical or chemical change?
2. How do you know?
3. What state of matter is Station 4?
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| **Station 5- Rusty Metal** Formula- 4Fe + 3O2 + 6H2O -> 4Fe(OH)31. Is this station a mixture or a pure substance?
2. How do you know?
3. When the metal rusted was that a physical or

 chemical change?1. What state of matter is Station 5?
 | **Station 6- Salt Water** Formula- NaCl + H2O1. Is the salt a mixture or pure substance?
2. Is the salt water a heterogeneous or homogeneous mixture?
3. What state of matter is Station 6?
4. How can we get the salt back?
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| **Station 7- One round ball (use your imagination)**1. Is this an example of an atom or molecule?
2. How do you know?
3. What state of matter is the ball at Station 7?
 | **Station 8- Dried Pasta and Beans**1. Is the pasta and the beans an example of a mixture or pure substance?
2. Is the pasta and bean mixture a heterogeneous or homogeneous mixture?
3. What state of matter is Station 8?
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| **Station 9- Sand** Formula- **SiO2 + CaSiO4 + Al2O3**1. Is the sand an example of a mixture or pure substance?
2. Is the sand a heterogeneous or homogeneous mixture?
3. What state of matter is Station 9?
 | **Station 10- Sugar Water** Formula- H2O + C12H22O111. Is the sugar water an example of a mixture or pure substance?
2. Is the sugar water a heterogeneous or homogeneous mixture?
3. What state of matter is Station 10?
4. How can I turn water into a gas?
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| ***Extra Credit*** **Station 11- Melted Crayons**1. Are these melted crayons a mixture or pure substance?
2. Is this an example of a chemical or physical change?
3. What could I do with the crayons to make a chemical change occur?
 | ***Extra Credit*** **Station 12- Air in a Balloon** Formula- 021. Is the air inside of balloon a mixture or a pure substance?
2. Is the air inside an element or a molecule?
3. What state of matter is the air?
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13) Look at the above Stations with “formulas”, using your textbook’s periodic table identify what each chemical abbreviation stands for. (Example- Mg= Magnesium) \*\*Each time you see a capital letter it represents a new element.