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| 1 | Ion  Vocabulary – Bonding #1 | Atom or group of atoms that has a positive or negative charge because it has lost or gained electrons |  |
| 2 | Ionic Bond | Chemical bond resulting from the transfer of electrons from one bonding atom to another |  |
| 3 | Covalent Bond | Chemical bond resulting from the sharing of electrons between two bonding atoms |  |
| 4 | Octet Rule | Rule that states that atoms tend to gain, lose, or share electrons so that each atom has a full outermost energy level, which is typically 8 electrons (octet) |  |
| 5 | Chemical Bond | The attractive force between atoms that results in the formation of a new substance |  |
| 6 | Lewis Dot Diagram (Structure) | Type of structural formula that uses dots or dashes to indicate bonds |  |
| 7 | Valence electrons | Electron in the outermost energy level of an atom; for most atoms, it is available to be gained, lost, or shared in the formation of chemical bonds. |  |
| 8 | Anion | Negative ion |  |
| 9 | Oxidation Number | Number assigned to the atoms in a molecule that shows the distribution of electrons among bonded atoms; equal to the charge in ionic compounds and the charge assigned the atom according to electronegativity rules in covalent compounds; sum of oxidation numbers in a molecule is zero |  |
| 10 | Cation | Positive ion |  |
| 11 | Crystalline Solid | Solid in which the atoms, ions or molecules are in a highly ordered, repeating pattern called a crystal |  |
| 12 | Inert | Nonreactive |  |

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