Worksheet #3: Drawing Ionic Bonds

Remember: *Ionic bonds form between POSITIVE IONS and NEGATIVE IONS. Ionic bonding is when one of the atoms is donating an electron(s) (the cation) and one of atoms is accepting an electron(s) (the anion). The electrons are not shared, the anion gains an electron(s) to achieve a full valence and the cation loses an electron(s) to achieve a full valence.*

Diagram the ionic bonding process from neutral atoms to ions showing the valence electrons and indicating with arrows the direction in which the electrons are going. Write your final answer in the box.

Ex: sodium nitride (Na_3N)

1. sodium chloride (NaCl)

2. barium oxide (BaO)

3. magnesium chloride ($MgCl_2$)

4. calcium chloride ($CaCl_2$)

5. potassium fluoride (KF)

6. sodium oxide (Na_2O)

7. aluminum chloride ($AlCl_3$)

8. rubidium oxide (Rb_2O)